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11: Section 1: Describing

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Formulas from Compound Names
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Governing How Compounds are

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Formed Naming Ionic Page
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Lewis Diagrams Made Easy:
How to Draw Lewis Dot
Structures ~~Common name,~~
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~~some chemical compounds.~~
Valence Electrons and the

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3: Naming and Writing
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1: Naming Ions Empirical

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Chapter 9: Chemical Names

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Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or

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F). Determined by the amount
of electrons lost or gained.

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To name a binary molecular
compound, use the following

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guidelines: 1. Write the names of the elements in the order listed in the formula
2. Use prefixes appropriately to indicate the number of each kind of atom
3. End the name of the second element with the

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suffix -ide Answers Page

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Names and Formulas
Flashcards ...~~

First, confirm that the
compound is a binary
molecular compound. Name the

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Formulas Answers Page

elements in the order listed in the formula. Use prefixes to indicate the number of each kind of atom. Omit the prefix -mono when the formula contains only one atom of the first element in the name. The suffix of the

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name of the second element
is -ide.

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1. when the name of the
anion (X) ends in -ide the

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acid name begins with the prefix hydro-. The stem of the anion has the suffix -ic and is followed by the word acid (ex. HCl= hydrochloric acid) 2. when the anion name ends in -ite the acid name is the stem of the anion

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with the suffix -ous
followed by the word acid
(ex. H_2SO_3 -- X=sulfite =
sulfurous acid)

~~CHAPTER 9 Chemical Names And
Formulas Test Review ...~~

Writing the Names and

Page 24/67

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Formulas for Bases. A base is a compound that produces OH^- in water. When naming a base, you name it like any other compound that starts with a regular or transition metal. Ex: NaOH = sodium hydroxide. When writing the

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Formula for a base, remember
to balance charges. Ex:

magnesium hydroxide = $\text{Mg}(\text{OH})$

2

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Chapter 9 Chemical Names and

Formulas 83 SECTION 9.3

NAMING AND WRITING FORMULAS

FOR MOLECULAR COMPOUNDS

(pages 268–270) This section
explains the rules for

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naming and writing formulas
for binary molecular
compounds. Naming Binary
Molecular Compounds (pages
268–269) 1. Circle the
letter of the type(s) of
elements that form binary
molecular compounds.

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~~Name Date Class CHEMICAL
221 NAMES AND FORMULAS 9~~

SECTION 9.2 NAMING AND
WRITING FORMULAS FOR IONIC
COMPOUNDS I. Write the
formulas for these binary
ionic compounds. c.

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potassium iodide 2. write
the formulas for the
compounds formed c. b. f. 3.

Name the following binary
ionic compounds. sodium
sulfide g. CuCl : h. SnCl_2 .

a. MnO_2 c. d. SrBr_2 221

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281 CHAPTER 9 Assessment 42.
a. 2- b. 1+ c. 1- d. 3+ 43.
a. 2+ b. 2+ c. 3+ d. 1+ 44.

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a. barium ion b. iodide ion
c. silver ion d. mercury(II)
ion 45. cyanide, CN^- and
hydroxide, OH^- 46. a.
hydroxide ion b. lead(IV)
ion c. sulfate ion d. oxide
ion 47. zero; A compound is
electrically neutral. 48.

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The symbols for the cation
and

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Write. Spell. Test. PLAY.

Match. Gravity. Created by.

bibiloni1901. Terms in this
set (20) Monatomic Ion.

Consists of a single atom
with a positive or negative
charge resulting from the
loss of gain of one or more

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Activities; Start over; Page

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Help; Use these activities

to help yourself to learn

the vocabulary and major

topics presented in this

chapter. A B; anion: any

atom or group of atoms that

has a negative charge:

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Chapter 9 Practice Test -
Naming and Writing Chemical
Formulas Answer Section

MATCHING 1. ANS: A PTS: 1

DIF: L1 REF: p. 253 OBJ:

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9.1.1 Identify the charges of monatomic ions by using the periodic table, and name the ions. STA: Ch.5.a 2.ANS: H PTS: 1 DIF: L1 REF: p. 254

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- 9.1 Naming Ions - 9.1

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Learning the fundamentals of
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chemical concepts and
problem solving. They'll
learn how to apply concepts
with the help of worked out
examples. In addition,
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and conceptual questions
checks brings together the

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useful to all readers
working in the chemical and
process industries and for
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safety and health

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