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## Answers: Moles and Stoichiometry

Practice Problems 1) How many moles of sodium atoms correspond to

$1.56 \times 10^{21}$  atoms of sodium?  $1.56 \times$

$10^{21}$  atoms Na  $\times$  1 mol Na =  $2.59 \times 10$

3 mol Na  $236.022 \times 10$  atoms Na 2)

Determine the mass in grams of each

of the following: a. 1.35 mol of Fe 1.35

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mol Fe x 55.845 g Fe = 75.4 g Fe 1  
mol Fe b. 24.5 mol O

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Answers: Moles and Stoichiometry  
Practice Problems

Moles and stoichiometry practice  
problems (from Chapter 3 in Brady,

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Russell, and Holum (Is Chemistry,  
Matter and its Changes, 3rdEd.)

Concept of mole/molar ratio

- 1) How many moles of sodium atoms correspond to  $1.56 \times 10^{21}$  atoms of sodium?
- 2) How many moles of Al atoms are needed to combine with 1.58 mol of O atoms to make

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aluminum oxide,  $\text{Al}_2\text{O}_3$ ? 3) How many moles of Al are in 2.16 mol of  $\text{Al}_2\text{O}_3$ ? 4) Aluminum sulfate,  $\text{Al}_2(\text{SO}_4)_3$ , is a compound used in sewage treatment plants. a.

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Problems (from Chapter 3...

Practice converting moles to grams, and from grams to moles when given the molecular weight. Practice converting moles to grams, and from grams to moles when given the molecular weight. If you're seeing this message, it means we're having



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Moles And Stoichiometry Practice  
Problems Answers | hsm1 ...

$x = 3.00$  mol of  $H_2$  was consumed.

Notice that the above solution used the answer from example #5. The solution below uses the information given in the original problem: Solution #2: The  $H_2 / H_2O$  ratio of 2/2 could

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Problems/Answers  
have been used also. In that case, the ratio from the problem would have been 3.00 over  $x$ , since you were now using the water data and not the oxygen data.

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Unit 4 Moles and Stoichiometry Mole  
Calculation Worksheet Answer Key

What are the molecular weights of the  
following compounds? 1) NaOH  $23 + 16 + 1 = 40.1$  grams  
2)  $H_3PO_4$   $3 + 31 + 64 = 98.0$  grams  
3)  $H_2O$   $2 + 16 = 18.0$  grams  
4)  $Mn_2Se_7$   $663.0$  grams

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5)  $\text{MgCl}_2$  95.3 grams 6)  $(\text{NH}_4)_2\text{SO}_4$  4  
132.1 grams Solve any 15 of the  
following: ...

---

Mole to Grams, Grams to Moles  
Conversions Worksheet  
Practice Problems: Stoichiometry.



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Balance the following chemical reactions: Hint a.  $\text{CO} + \text{O}_2 \rightarrow \text{CO}_2$  b.  $\text{KNO}_3 \rightarrow \text{KNO}_2 + \text{O}_2$  c.  $\text{O}_3 \rightarrow \text{O}_2$  d.  $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$  e.  $\text{CH}_3\text{NH}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{N}_2$  Hint f.  $\text{Cr}(\text{OH})_3 + \text{HClO}_4 \rightarrow \text{Cr}(\text{ClO}_4)_3 + \text{H}_2\text{O}$  Write the balanced chemical equations of each reaction:

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## Practice Problems: Stoichiometry

While the mole ratio is ever-present in all stoichiometry calculations, amounts of substances in the laboratory are most often measured by mass.

Therefore, we need to use mole-mass

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Problems Answers  
Calculations in combination with mole ratios to solve several different types of mass-based stoichiometry problems.

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12.3: Mass-Mole and Mole-Mass  
Stoichiometry - Chemistry ...

*Page 27/68*

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Determine the amount (in moles) of a product from a given amount of one reactant. Determine the amount (in moles) of a product from a given amount of one reactant. If you're seeing this message, it means we're having trouble loading external resources on our website. ... Practice:

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Ideal stoichiometry (practice) | Khan  
Academy

Stoichiometry I: Mole-Mole Problems \*

Description/Instructions ; To solve  
mole-mole problems requires a

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balanced chemical equation and a mole ratio. Use the coefficients from the balanced equation and multiply it by the appropriate mole ratio to get an answer. This quiz will cover simple mole-mole problems. You will need a calculator.

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Stoichiometry : Stoichiometry I: Mole-  
Mole Problems Quiz

Stoichiometry example problem 1.

Stoichiometry example problem 2.

Practice: Ideal stoichiometry. Practice:  
Converting moles and mass. This is  
the currently selected item. Next

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12/2020  
Answers: Moles and Stoichiometry  
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is ever-present in all stoichiometry  
calculations, amounts of substances in  
the laboratory are most often  
measured by mass. Therefore, we  
need to use mole-mass calculations in

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combination with mole ratios to solve  
several different types of mass-based  
stoichiometry problems.

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20 Then do some stoichiometry using

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Easy math 16 g of methane (MM = 16) is 1 mole and 1 mole of methane will produce 1 mole of  $\text{CO}_2 = 44 \text{ g}$ , and 2 moles of  $\text{H}_2\text{O}$  which is 36 g for a total of 80 g

4. d Balance:  $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

5. d Balance:  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

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Practice Test Ch 3 Stoichiometry

Name Per

5. A comprehensive problem on reaction stoichiometry: mole ratio, limiting reactant, percent yield and amount of reactants needed. Aspirin (acetyl salicylic acid) is widely used to

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Problems Answers  
treat pain, fever, and inflammation.

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out <http://socratic.org/chemistry> Lots  
and lots and lots of practice problems

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with mole ratios. This is the first step  
in...

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This chemistry video tutorial provides  
a basic introduction into stoichiometry.

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It contains mole to mole conversions,  
grams to grams and mole to gram  
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Stoichiometry Basic Introduction, Mole  
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Answers: Moles and Stoichiometry

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$1.56 \times 10^{21}$  atoms of sodium?  $1.56 \times 10^{21}$  atoms Na  $\times$   $1 \text{ mol Na} = 2.59 \times 10^{-3} \text{ mol Na}$   $236.022 \times 10$  atoms Na 2)

Determine the mass in grams of each of the following: a. 1.35 mol of Fe  $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$  1



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mol Fe b. 24.5 mol O 24.5 mols O

A Visual Analogy Guide to Chemistry  
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